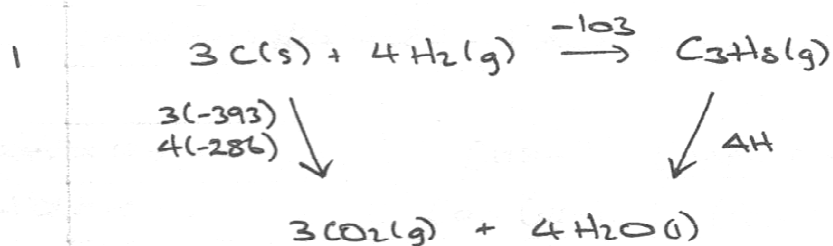


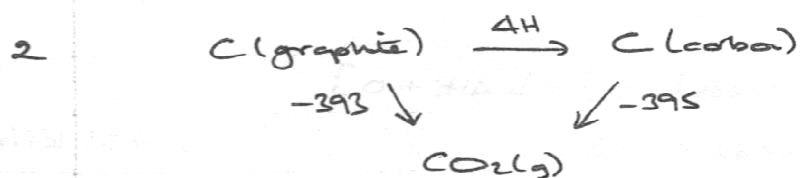


HESS'S LAW 2 - COMBUSTION



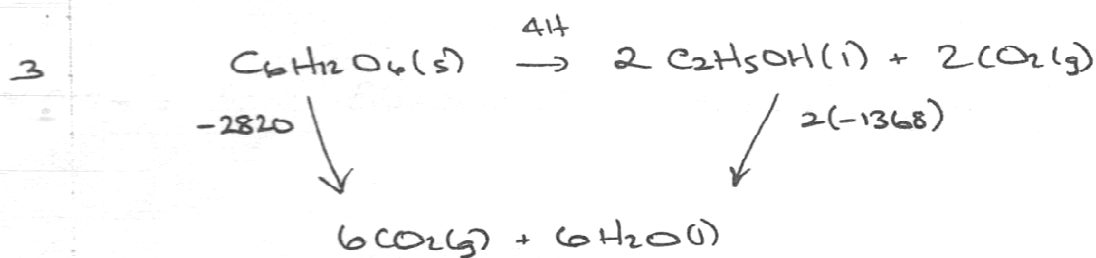
$$\Delta H - 103 = 3(-393) + 4(-286)$$

$$\Delta H = \underline{-2220 \text{ kJ/mol}}$$



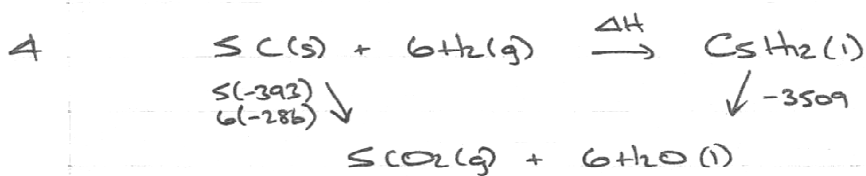
$$\Delta H - 395 = -393$$

$$\Delta H = \underline{+2 \text{ kJ/mol}}$$



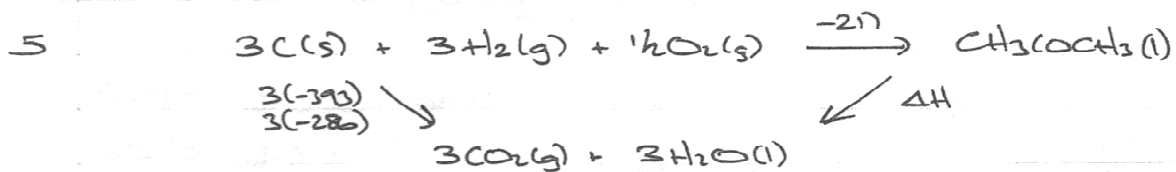
$$\Delta H + 2(-1368) = -2820$$

$$\Delta H = \underline{-84 \text{ kJ/mol}}$$



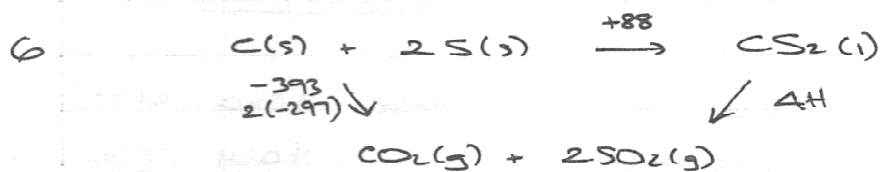
$$\Delta H - 3509 = 5(-393) + 6(-286)$$

$$\Delta H = \underline{-172 \text{ kJ/mol}}$$



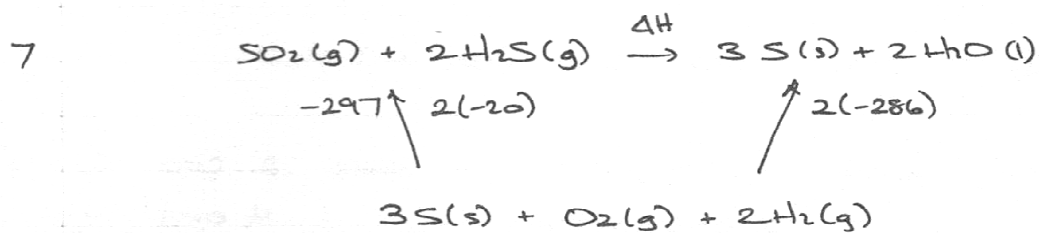
$$\Delta H - 217 = 3(-393) + 3(-286)$$

$$\Delta H = \underline{-1820 \text{ kJ/mol}}$$



$$\Delta H + 88 = -393 + 2(-297)$$

$$\Delta H = \underline{-1075 \text{ kJ/mol}}$$



$$\Delta H - 297 + 2(-20) = 2(-286)$$

$$\Delta H = \underline{-235 \text{ kJ/mol}}$$