2

HESS'S LAW 2 - COMBUSTION

$$3 C(s) + 4 Hz(g) \xrightarrow{-103} C3Hs(g)$$

 $3(-393)$
 $4(-286)$
 $3(02(g) + 4 Hz(0))$

C(graphia)
$$\xrightarrow{AH}$$
 C(coroa)
-393 \ \(\sigma -395 \)

CO2(3)

$$\Delta H = \frac{+2 \, kT \, lmal}{}$$

```
SC(5) + 6+2(9) -> CSH2(1)
4
         S(-393) V
              S(02(q) + 6+20(1)
         44-3509 = 5(-393) + 6(-286)
         4H = - 172 KJ/mol
         3(5) + 3+2(g) + 1hOz(g) -21) CH3(OCH3(1)
3(-393) AH
5
          AH-217 = 3(-393) + 3(-286)
        AH = -1820 KJ/MD1
         C(5) + 25(3)
         -393
2(-297)
         1 (Oz(5) + 250z(3)
          AH + 88 = -393 + 2(-297)
          AH = -1075 KU/mal
          502(g) + 2H25(g) -> 35(s)+2HO(1)
 7
       35(s) + O2(g) + 2H2(g)
         AH-297+2(-20) = 2(-286)
       AH = -235 KJ/NEI
```